

H2o renewable energy solutions



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Why can you only add oleum to water to get SO₂ (g)

Why can you only add oleum to water to get SO₂ (g) industrially instead of adding SO₃ (g) to H₂O (l) immediately?

What are the names of the two molecules plants need to

Two molecules that plants need to absorb to perform photosynthesis are CO₂ and H₂O.



Question #545aa

$HA(aq) + OH^-(aq) \rightarrow A^-(aq) + H_2O(l)$ It's important to notice that you have 1:1 mole ratios between the weak acid, the strong base, and the conjugate base A⁻. This means that for every one mole of weak

H2O Renewables , Manchester Heat Pumps

We are specialists in the design, supply & installation for eco-friendly energy solutions such as heat pumps, solar panels, Tesla Powerwall 3 battery storage, biomass and lpg in Manchester and the



Why is "H"₂ not very soluble in water? , Socratic



Question #9b627 + Example

Changes in state are physical changes. Let's take water as an example. H₂O. In the gaseous state, we sometimes call it steam or water vapor; it's chemical formula is still H₂O. In the liquid state and in the



Questions asked by

CH₄ + 2 O₂ -> CO₂ + 2 H₂O? The specific heat capacity of platinum is 0.032 cal/g.K. How to calculate the heat, in joules, necessary to raise the temperature of a sample of platinum weighing 20.0 g from



The trade in energy sucks for H₂ (but not so much for water), so the H₂?H₂O interaction is highly unfavored for H₂, compared to the H₂?H₂ and H₂O?H₂O interactions. So, H₂ is very



Writing word equations as chemical equations and balance

magnesium oxide = MgO Hydrochloric acid = HCl
 Magnesium Chloride = MgCl₂ Water = H₂O MgO
 + 2HCl rarr MgCl₂ + H₂O Beryllium chloride =
 BeCl₂ Silver nitrate = AgNO₃ Beryllium



An aqueous solution contains 32.7% KCl. How many grams of

An aqueous solution contains 32.7% KCl. How many grams of KCl are contained in 100 g of this solution? Chemistry

Using ammonia explain the difference between strong and weak

Well, consider the reaction of ammonia in aqueous solution . $\text{NH}_3(\text{g}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{NH}_4^+ + \text{OH}^-$ Now all acid-base behaviour is modified by the identity of the



Complete combustion of C14H18O7 to CO2 and H2O. You have 315

Complete combustion of C14H18O7 to CO2 and H2O. You have 315.0000 grams of C14H18O7. How many grams of O2 are required? How many grams of CO2 are generated? How many grams of H2O

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